

Activity series of the elements

	Strongly Reducing	$\text{Li}(s) \rightarrow \text{Li}^+(aq) + e^-$ $\text{K}(s) \rightarrow \text{K}^+(aq) + e^-$ $\text{Ba}(s) \rightarrow \text{Ba}^{2+}(aq) + 2e^-$ $\text{Ca}(s) \rightarrow \text{Ca}^{2+}(aq) + 2e^-$ $\text{Na}(s) \rightarrow \text{Na}^+(aq) + e^-$		Weakly Oxidizing	React rapidly with $\text{H}^+(aq)$ or $\text{H}_2\text{O}(\ell)$ to release $\text{H}_2(g)$
		$\text{Mg}(s) \rightarrow \text{Mg}^{2+}(aq) + 2e^-$ $\text{Al}(s) \rightarrow \text{Al}^{3+}(aq) + 3e^-$ $\text{Mn}(s) \rightarrow \text{Mn}^{2+}(aq) + 2e^-$ $\text{Zn}(s) \rightarrow \text{Zn}^{2+}(aq) + 2e^-$ $\text{Cr}(s) \rightarrow \text{Cr}^{3+}(aq) + 3e^-$ $\text{Fe}(s) \rightarrow \text{Fe}^{2+}(aq) + 2e^-$			React with $\text{H}^+(aq)$ or steam to release $\text{H}_2(g)$
		$\text{Cd}(s) \rightarrow \text{Cd}^{2+}(aq) + 2e^-$ $\text{Co}(s) \rightarrow \text{Co}^{2+}(aq) + 2e^-$ $\text{Ni}(s) \rightarrow \text{Ni}^{2+}(aq) + 2e^-$ $\text{Sn}(s) \rightarrow \text{Sn}^{2+}(aq) + 2e^-$ $\text{Pb}(s) \rightarrow \text{Pb}^{2+}(aq) + 2e^-$			React with $\text{H}^+(aq)$ to release $\text{H}_2(g)$
		$\text{H}_2(g) \rightarrow 2\text{H}^+(aq) + 2e^-$			
		$\text{Cu}(s) \rightarrow \text{Cu}^{2+}(aq) + 2e^-$ $2\text{I}^-(aq) \rightarrow \text{I}_2(s) + 2e^-$ $\text{Ag}(s) \rightarrow \text{Ag}^+(aq) + e^-$ $\text{Hg}(\ell) \rightarrow \text{Hg}^{2+}(aq) + 2e^-$ $2\text{Br}^-(aq) \rightarrow \text{Br}_2(\ell) + 2e^-$ $\text{Pt}(s) \rightarrow \text{Pt}^{2+}(aq) + 2e^-$ $2\text{H}_2\text{O}(\ell) \rightarrow \text{O}_2(g) + 4\text{H}^+(aq) + 4e^-$ $2\text{Cl}^-(aq) \rightarrow \text{Cl}_2(g) + 2e^-$ $\text{Au}(s) \rightarrow \text{Au}^{3+}(aq) + 3e^-$ $2\text{F}^-(aq) \rightarrow \text{F}_2(g) + 2e^-$			Do not react with $\text{H}^+(aq)$ to release $\text{H}_2(g)$
	Weakly Reducing			Strongly Oxidizing	